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Worksheet Answers

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$q = m \times c \times (t. = 25 \text{ g} \times 0.5050 \text{ J/g}^\circ\text{C} \times (15.6 - 10.5^\circ\text{C}) = 64.39 \text{ J}$ . A piece of aluminum with a mass of 50g and an initial temperature of 90oC is placed into 100mL of water at a temperature of 25oC. The temperature of water rises to

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31.30C. Determine the specific heat capacity of aluminum.

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Definitions:-Endothermic:-Exothermic:-Enthalpy:-Law of Conservation of Energy-Heat of Reaction:-Heat of Fusion-Specific Heat. 2. Explain why water is used in a calorimeter? What unique properties does water have? Calculations:

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#### AP CHEMISTRY REVIEW WORKSHEET

(Unit 4 - Thermochemistry) 1.  $3 \text{ SO}_2 (\text{g}) + 2 \text{ HNO}_3 (\text{aq}) + 2 \text{ H}_2\text{O} (\text{l}) \rightarrow 3 \text{ H}_2\text{SO}_4 (\text{aq}) + 2 \text{ NO} (\text{g})$  a. Calculate  $\Delta H$  for the above reaction. Based on the enthalpy change, determine whether the reaction



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is endothermic or exothermic and whether it would be thermodynamically favored or unfavored.

**AP CHEMISTRY REVIEW WORKSHEET  
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(handouts) Chapter 18 - reaction rates

(handouts) Chapter 19 - acids, bases,

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Chemistry Basics; Forming Crystals;  
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## **Thermochemistry Worksheets - Printable Worksheets**

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thermochemistry practice problems that are typical for a chapter on this topic. These problems can be used for homework or for test review. These problems cover: endothermic and exothermic reactions, heat gain and heat loss, heat of fusion, heat of crystallization, h.

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**Thermochemistry Worksheets & Teaching Resources | TpT**

a. calorimeter d. enthalpy b. calorie e.  
Match each item with the correct  
statement below. a. heat of reaction d.  
heat of fusion b. heat of formation e.  
Multiple Choice Identify the letter of the  
choice that best completes the  
statement or answers the question.

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Chemistry Site ... Schedule

## **Thermochemistry - Miss Shafer's Chemistry Site**

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True or False If the answer is false you must explain why 1 Chemical energy is a form of Thermochemistry Review.pdf - Thermochemistry Review...

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Thermochemistry Answers - Worksheet Number One. We will ignore any heats



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losses to the walls of the container and losses to the air. These is a typical position to take since, in a real experiment, both would have to be accounted for, making for much more complexity. 1.  $q = (20.0 \text{ g}) (20.0 \text{ }^\circ\text{C}) (2.02 \text{ J/g }^\circ\text{C})$ .

## **Thermochem WS #1 Answers -**

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AP CHEMISTRY REVIEW WORKSHEET

(Unit 4 - Thermochemistry) S + 2

HN03(aq) + 21-170 (l) 3 + 2 NO (g)

PERIOD: a D/Ö-/l a. Calculate AH for the above reaction. Based on the enthalpy change, determine whether the reaction is endothermic or exothermic and whether it would be thermodynamically

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favored or JPs unfavored. b. c. 2]

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California State University Affordable Learning Solutions Program, and Merlot. We also acknowledge previous National Science Foundation support under grant numbers 1246120, 1525057, and 1413739.

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The amount of heat needed to increase the temperature of 1g of a substance 1 degree Celsius; also called specific heat capacity.

**Thermochemistry Vocabulary  
Flashcards | Quizlet**

Ch 17 Thermochemistry Practice Test  
Matching Match each item with the

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correct statement below. a. calorimeter  
d. enthalpy b. calorie e. specific heat c.  
joule f. heat capacity \_\_\_\_ 1. quantity of  
heat needed to raise the temperature of  
1 g of water by  $1^{\circ}\text{C}$  \_\_\_\_ 2. SI unit of  
energy \_\_\_\_ 3.

## **Ch 17 Thermochemistry Practice Test**

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